


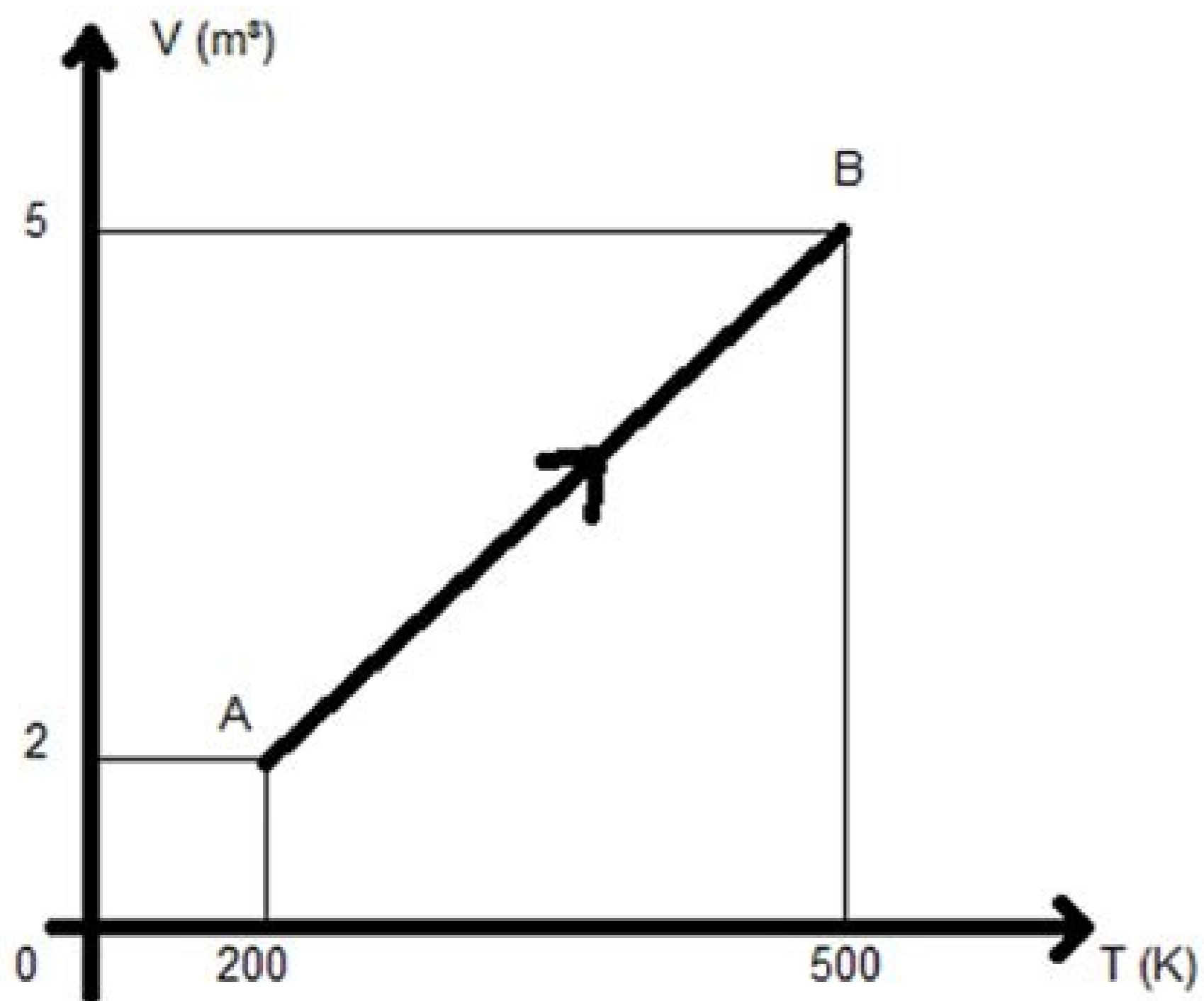
**Cal/mol para j/mol**

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**Next**

- c) A variação de energia interna sofrida pelo gás.  
(Considere  $R = 8,31 \text{ J/mol} \cdot \text{K}$ )

A quantidade de 3 mols de um gás ideal monoatômico sofre a expansão isobárica AB representada no gráfico. Sendo o calor molar sob pressão constante desse gás  $C = 5 \text{ cal/mol} \cdot \text{K}$  e adotando  $R = 8,31 \text{ J/mol} \cdot \text{K}$ , determine:



- a) A pressão sob a qual o gás se expande;  
b) A quantidade de calor recebida pelo gás;  
c) O trabalho eu o gás realiza na expansão;  
d) A variação de energia interna sofrida pelo gás.  
(Considere  $1 \text{ cal} = 4,18 \text{ J}$ )

Na questão anterior, se o aquecimento de 200 K a 500 K fosse isocórico, qual seria a quantidade de calor recebida pelo gás? Considere  $R = 2 \text{ cal/mol} \cdot \text{K}$ .

1/1

$$\bar{C} = \frac{\text{J}}{\text{mol} \cdot ^\circ\text{C}}$$

INDICE

Leyes de los gases

**Ecuación general de los gases ideales**

Combinación de las tres leyes:

$$V = \frac{k'k''k'''nT}{P} = \frac{RnT}{P} \quad \rightarrow \quad \text{Ley de los gases ideales: } PV = nRT$$

R se calcula para:

$n = 1 \text{ mol}$ $P = 1 \text{ atm}$ $V = 22,4 \text{ l}$ $T = 273 \text{ K}$	$R = 0,082 \text{ atm L/mol K}$	$\frac{P \cdot V}{T} = \frac{P' \cdot V'}{T'}$
$R = 8,31 \text{ J/mol K} = 1,987 \text{ cal/mol K}$		

$$\frac{p_1 \cdot V_1}{T_1} = \frac{p_2 \cdot V_2}{T_2}$$

$$p \cdot V = n \cdot R \cdot T$$

Something went wrong. Wait a minute and try again. Latest update: May 24th, 2021 `124Whach6; 128? agree with organic chemicals using kilocal (kcal) instead of kilojoule s (kJ)? The metric system of the International System (SI) was adopted for the first time in France in the 1791 shortly after the French Revolution (one of their other innovations adopted by the revolutionaries the week of ten days did not survive). D 160? I am Canadian, so I grew up entirely within the metric system, and since the units are the de rigneur of measurement in science, I just assumed that all scientists used it. When I started studying organic chemistry etc 160; It was like finding a den of shameful heretics operating right under the noses of the Inquisition. All the values of strength points and related energies that I learned and memorized in kilojoule/mol had been returned to me in kilocal. Being Canadian (and a new doctorate)19; 160? I kindly withheld my criticism of this barbaric practice until a more opportune moment, while maintaining my estimation that IT IS 160; was the most correct system. [Only to refresh your memory, one gram of water per degree Celsius is a unit that measures energy. Itis the energy needed to increase the temperature of the first gram of water per degree Celsius. Even though it is active; it is still u se d as a food energy unit, the calorie has been replaced by u Giostra Systeme Internationale (J). After a few weeks of listening to my colleagues talking in kilocali, I started arriving at dawn which was only my research group that did this. It was the entire community of organic chemists in magazines, conferences, casual conversation. And gradually, I started using it, too. D 160? At first I felt a little bit frisky to get away from my education system. Month by month, however, my resistance to the use of kcal began to crumble the more I used it, the more I realized it was actually a unit is more practical for the organic the organic Chem Gen we learn to look at the heats of formation of whole molecules, which can be quite large numbers. On the contrary, when organic chemists speak of energies, they usually speak of energies and conformations of dissociations. Since the variety of molecules in organic chemistry is so enormous, we often do not have the luxury of being able to work with values from tables. Therefore, thinking about energy changes in a reaction, we are often forced to base our thinking on estimating the value of the parts rather than the whole. 3What is that C-O bond worth? What is the cost of replacing it with bromide? What's the energy for that ring spin? What is the contribution of that hydrogen bond?â. It's like going from work at a bounty store and one day you're looking at a car and thinking, "What's that bumper worth?" So why do organic chemists use kcal? I can think of four reasons. 1. The C-H binding force in alkanes of 100 kcal/mol is a convenient mental anchor. Organic chemists think a lot in terms of binding force. The most common bond you will find in organic chemistry is the C-H bond, which has a binding force of about 100 kcal/mol (note: this value can vary greatly depending on the carbon substitution, but for simple alkanes such as methane and ethane, the C-H bond is very close to this value). What is this value in kJ/mol? 472 kJ/mol. Guess which number is the easiest to remember? There's just something instinctive about working on a 0-100 scale. 2. Low energy phenomena also fit well on the kcal scale. For organic chemistry, kcal/mole is a kind of "goldilocks" metric, as even 1 kcal/mole is significant, but not too significant. For example, 1 kcal/mol is not even close to being enough energy to break a bond, but it is still enough to cause an 83:17 ratio of products in equilibrium. Going a little longer High, hydrogen ties have strengths of anywhere from 2-7 kcal / mol. Long the same When you get to know the conformations it is said that the barrier to the rotation of the titanium (from the "eclipsed" form to the "scaled" form) is about ten kJ/mol. Converted to kcal, So on the one hand you have the strength of the C-H to 100 bond and on the other you have hydrogen bonds and conformational changes from about 2-7 kcal/mol. The scale incorporates them both very well. 3. In general, the smaller numbers are easier to handle. One thing of the strength values of the bond you will see is that they vary a lot. Depending on where you look you can see values of 78-82 kcal/mol for the resistance of the C-C bond. Facing a 4-uncertainty seems much simpler than the kJ equivalent, which would be 320-340 kJ/mol. Maybe it's just me, but I like dealing with smaller numbers. As a commentator said when I asked Chemistry Reddit about this a few days ago, kcals make my errors seem smaller. 4. In fact, according to Wikipedia, the metric system was developed in England and introduced in France by Benjamin Franklin. Is that something Ben Franklin wasn't involved in? Bottom line: Organic chemists find kcal/mol much cheaper to use as an energy unit. Personally, I think and speak in kcal/mol, but I don't want to increase the confusion that people already experience for organic chemistry, so I try to keep things in kJ/mol here, but if I step back in energy measure I find it more comfortable to work with, that's why ©. why ©.

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